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provided. 1. The reaction for the decomposition of hydrogen peroxide is  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ .

List three ways to speed up the rate of decomposition. For each one, briefly explain why it is effective,

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based on collision theory.

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